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(54) Title: A METHOD AND APPARATUS FOR THE PRODUCTION OF METAL COMPOUNDS

(57) Abstract: The present invention relates to a method and apparatus for the production of titanium alloys and titanium-aluminium inter-metallic compounds and alloys. Starting from a precursor material including titanium subchloride (titanium trichloride or titanium dichloride), the precursor material is reduced by aluminium to produce titanium-aluminium intermetallic complexes or alloys and aluminium chloride which is driven away from the reaction zone so as to favour the forward reaction and the production of the titanium-aluminium compounds. Starting from a precursor material of titanium subchloride avoids the problems associated with starting from titanium metal (which is expensive to produce) or titanium tetrachloride (a reaction very difficult to control), and results in the production of powdered forms of titanium-aluminium compounds with controllable composition.

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**A METHOD AND APPARATUS FOR THE PRODUCTION OF METAL COMPOUNDS**

Field of the Invention

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The present invention relates to a method and apparatus for the production of metal and metal compounds and, particularly, but not exclusively, to a method and apparatus for production of titanium-based alloys and 10 intermetallic complexes, and more particularly, but not exclusively, to a method and apparatus for the production of titanium-aluminium based alloys and intermetallic complexes, and more particularly, but not exclusively, to a method and apparatus for the production of titanium-15 aluminium based alloys and intermetallic complexes in a powder form.

Background of the Invention

20 Titanium-aluminium alloys and inter-metallic compounds (generically termed herein "titanium-aluminium compounds") are very valuable materials. However, they are difficult and expensive to prepare, particularly in the preferred powder form. This expense of preparation limits wide use of these materials, even though they have 25 highly desirable properties for use in automotive, aerospace and other industries.

Titanium minerals are found in nature in the form of a very stable oxide ( $TiO_2$ ). Common processes for the 30 production of titanium are the Kroll process and the Hunter process. The Kroll process requires the use of magnesium as a reducing agent to reduce  $TiCl_4$  (prepared from the oxide by a pre-process of chlorination) to produce the Ti metal. The Hunter process requires the use 35 of sodium as the reducing agent. Because  $TiCl_4$  is still thermodynamically stable, highly reactive reducing agents such as magnesium or sodium are required to produce

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titanium metal out of  $TiCl_4$ . Such highly reactive reducing agents are difficult and expensive to handle. As the magnesium chlorides in the case of the Kroll process are stable up to temperatures in excess of 1300K, the product 5 is often in the form of a Ti sponge mixed with  $MgCl_2$ , and remnants of Mg and  $TiCl_2$ . To obtain pure Ti, the product requires extensive post-processing, including washing and melting in a vacuum arc furnace to remove all impurities. This contributes to the present high cost of the 10 production of titanium.

In the known technologies for production of titanium alloys such as Ti-Al-V, and intermetallic compounds such as  $Ti_3Al$ ,  $TiAl$ ,  $TiAl_3$ , Ti-Al-(Cr, Nb, Mo, etc) and alloys based on these compounds, appropriate amounts of sponges, 15 ingots or powders of the metals which comprise these alloys are milled or melted together and annealed, hence adding to the production cost, particularly as it is necessary to obtain the metals first which, as discussed, in the case of titanium, involves considerable expense. 20 For production of a powder of these titanium alloys and intermetallic compounds, further processing is usually required, adding to the already high production cost.

Prior Al-based processes for manufacturing of Ti-Al compounds include starting from Al powder and Ti powder 25 (references: (I.Lu, M.O. Lai and F.H. Froes, *Journal of Metals*, February 2002, p62) and (N. Bertolino et al., *Intermetallics*, Vol 11, 2003, p 41) and reduction of  $TiCl_4$  with  $AlCl$  (US patent application US2002/0184971 A1)). For the first process, the starting materials are Al and Ti 30 powders, the powders usually being mechanically milled to make a uniform mixture followed by heating in a furnace. The resulting materials are at best in the form of solid lumps and this process is usually unable to produce fine powder. Furthermore, the resulting compounds often 35 require heat treatment to produce the required material properties. For the second process, Al metal is heated in the presence of chlorine at temperature around 1200C to

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produce gaseous AlCl that is then reacted with TiCl<sub>4</sub> in the gas phase to produce powders of titanium aluminides. Both these processes are quite complex and costly to operate.

It is also known to perform direct reduction of TiCl<sub>4</sub> with aluminium. However, this results in the production of an uncontrollable composition of compounds and production of a single phase material such as TiAl has not been achieved (see in particular S J Gerdemann & D E Alman, page 3341 in *Gamma Titanium Alumini* 1999, edited by Kim, Dimiduk & Loretto, The Minerals Metals and Materials Society USA).

Over the past several decades, there have been extensive attempts made to replace the existing Kroll and Hunter technologies using techniques such as electrowinning, plasma-hydrogen and also aluminothermic reduction.

The use of hydrogen plasma for the reduction of titanium chloride in a plasma atmosphere is difficult due to unfavourable thermodynamic characteristics, since chlorine preferably reacts with titanium in the reverse reaction to produce titanium chlorides, hence degrading the quality of the produced Ti powder and limiting the efficiency of the method. In a process disclosed in US Patent 5,935,293, a fast quench reactor was used to cool down the plasma in order to prevent recombination processes leading to formation of titanium chlorides. According to the description in US Patent 5,935,293, the process is highly energy expensive relative to the existing Kroll technology.

In another process (G.Z. Chen, D.J. Fray and T.W. Farthing, *Nature*, Vol 407, (2000), 361), Chen et al. made titanium sponge directly from the oxide by reduction in a molten calcium chloride salt. Oxygen from the titanium oxide recombines with carbon at an anode to form CO<sub>2</sub>. However, the composition of the resulting sponge-like titanium product produced corresponds to the composition of the starting minerals. The process is still under

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development and is yet to be demonstrated on an industrial scale.

Attempts have been made to use aluminium as a reducing agent for  $TiCl_4$  in plasma systems. For reduction 5 of  $TiCl_4$  using aluminium, the products are in the form of solid phase titanium-aluminium intermetallic compounds mixed with aluminium chloride and some residual titanium dichloride. A description of various attempts using aluminium together with a description of the 10 thermodynamics of the process are given by Murphy and Bing (*High Temp. Chem. Processes*, Vol 3, 365-374, 1994). Because of difficulties associated with gas phase reactions it has not been possible to produce titanium and/or titanium-aluminium compounds by direct 15 aluminothermic reduction of titanium chlorides.

#### Summary of the Invention

In accordance with a first aspect, the present 20 invention provides a method for the production of titanium compounds, comprising the steps of mixing aluminium with a precursor material including titanium subchloride, and heating the mixture, to form aluminium chlorides and titanium compounds.

When the term titanium subchloride is used, it can refer to titanium trichloride  $TiCl_3$  and/or titanium dichloride  $TiCl_2$  or other combinations of titanium and chloride excluding  $TiCl_4$  which is referred to herein as titanium chloride.

When the term titanium compound is used, it can refer 30 to titanium alloys and/or titanium/metal intermetallic compounds. In one preferred form which is referred to herein, the titanium compounds include titanium-aluminium alloys and/or titanium-aluminium intermetallic compounds.

In one embodiment the method can also comprise the 35 step of arranging reaction conditions to favour a forward reaction to form the aluminium chlorides and titanium

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compounds. The step of arranging the reaction conditions can involve driving aluminium chlorides away from a reaction zone where the aluminium and precursor material are reacting. In one form, the step of arranging the  
5 reaction conditions can include a continuous removal of aluminium chlorides from the reaction zone.

The method can also comprise the step of preparing the titanium subchloride precursor material from titanium chloride ( $TiCl_4$ ). In one embodiment, the step of preparing  
10 titanium subchloride is carried out by the reduction of  $TiCl_4$  using aluminium. In an alternative embodiment, the step of preparing titanium subchloride is carried out by heating  $TiCl_4$  in a plasma of an argon-hydrogen mixture. In one embodiment,  $TiCl_3$  is prepared by the reduction of  
15  $TiCl_4$ , although this reaction may also form titanium dichloride  $TiCl_2$ .

In either embodiment, the method can also comprise the further step of recycling at least some of the aluminium chloride formed, and utilising the aluminium  
20 chloride to produce  $TiCl_4$ . The aluminium chloride can be used to reduce titanium oxide to produce  $TiCl_4$ .

In any of the embodiments of the method mentioned, aluminium oxide can be produced by the reduction of titanium oxide and the aluminium oxide electrolysed to  
25 produce aluminium raw material for use in the steps of said method.

In one embodiment of the method, the aluminium chlorides are condensed away from the reaction zone at a temperature lower than that in the reaction zone. In one  
30 embodiment of the method, if titanium subchloride escapes the reaction zone it can be condensed at a temperature different to that in the reaction zone. Furthermore the condensed titanium subchloride can optionally be returned to the reaction zone.

35 In one form of the method, the precursor material can include vanadium subchloride, and a product of said method can be an alloy or intermetallic complex including

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titanium, aluminium and vanadium. This method can also include the steps of mixing the precursor material in appropriate proportions and carrying out the method to produce Ti-6Al-4V.

5 In a further form of the method, the precursor material can include zirconium subchloride, and a product of the method can be an alloy or intermetallic complex including titanium, aluminium, zirconium and vanadium.

In one form of the method, the precursor material can 10 include niobium halide and chromium halide, and a product of said method can be an alloy or intermetallic complex including titanium, aluminium, niobium and chromium. This method can also include the steps of mixing the precursor materials in appropriate proportions and carrying out the 15 method to produce Ti-48Al-2Nb-2Cr.

In further embodiments of the method, the precursor material can include a source of one or more elements selected from the group comprising chromium (Cr), niobium (Nb), vanadium (V), zirconium (Zr), silicon (Si), boron 20 (B), molybdenum (Mo), tantalum (Ta) and carbon (C), and products of said method can include titanium-aluminium compounds which include one or more of these elements. The source of the element(s) can be a metal halide, a subhalide, a pure element or another compound which 25 includes the element. The products can also include one or more of an intermetallic compound, a titanium-(selected element)-alloy, and intermediate compounds.

In still further embodiments, the precursor materials may include other metals and/or metal halides, and the 30 products of said method may include titanium alloys or intermetallic compounds. For example the precursor material may include vanadium subchlorides, such as vanadium trichloride and/or vanadium dichloride, and the products may include one or more of a titanium-aluminium-vanadium alloy, an intermetallic compounds, and a 35 combination of a titanium-vanadium alloy and an intermetallic compound. The precursor materials may also

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include a source of other precursors containing a required alloy additive, depending upon the required end product.

In one form of the method, the aluminium can be added in the form of a powder having an approximate upper grain size of less than about 50 micrometres. In an alternative 5 form of the method, the aluminium can be in the form of a powder of an approximate upper grain size of greater than about 50 micrometres, and the method comprises the step of milling the aluminium powder and titanium subchloride to reduce the grain size of the aluminium powder in at least 10 one dimension. In yet another alternative form of the method, the aluminium can be in the form of flakes having a thickness in one dimension of less than about 50 micrometres. Using a fine aluminium powder is preferred, 15 although the relatively coarser aluminium powder or flakes is a cheaper raw material.

In any of the embodiments described, the method can also comprise the further step of adding a reagent to a product of the method to produce a further product.

20 In any of the embodiments described, the method can also include a pre-processing step of forming the titanium subchloride as precursor material.

The inventor has found that using a precursor material which includes titanium subchloride (preferably 25 titanium trichloride) gives a number of advantages. There are not the problems of different, uncontrollable phases which are involved in starting from titanium tetrachloride as a precursor. The composition of the end product is relatively controllable and depends on the ratios of the starting materials. The correct ratios of starting 30 materials are incorporated in the precursor materials to produce the appropriate proportions of components in the product.

The inventor believes that the new method enables a 35 cheaper and more controllable process for the production of titanium-aluminium compounds. It is not necessary to convert the raw titanium minerals to titanium metal, as in

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some of the prior art processes discussed earlier. In one embodiment of the present process, titanium oxide can be chlorinated using conventional technology to give titanium tetrachloride. This can then be reduced using aluminium or hydrogen to give titanium subchlorides (mainly titanium trichloride), which can then be used as the precursor material for the formation of the titanium-aluminium compounds.

It is possible to form Ti-6Al-4V using this process, which is one of the major titanium alloys used. It is also possible to form Ti-48Al-2Nb-2Cr. It is also possible to form other alloys such as Ti-Al-Nb-C, and Ti<sub>3</sub>Al based alloys.

A process in accordance with embodiments of the present invention described has the advantage that alloy powder is produced directly, with no further physical processing.

In accordance with a second aspect, the present invention provides a method of producing titanium-aluminium compounds, comprising the steps of:

- heating a precursor material of TiCl<sub>4</sub> in a plasma of an argon-hydrogen mixture to produce TiCl<sub>3</sub>;
- mixing aluminium with said TiCl<sub>3</sub>; and
- heating the resultant mixture to produce titanium-aluminium compounds and AlCl<sub>3</sub>.

In accordance with a third aspect, the present invention provides a method of producing titanium-aluminium compounds, comprising the steps of:

- heating a mixture of TiCl<sub>4</sub> and aluminium to form TiCl<sub>3</sub> and AlCl<sub>3</sub>;
- heating the mixture to a reaction zone temperature above 300C and providing for AlCl<sub>3</sub> to be evaporated from the reaction zone;
- adding further aluminium to the mixture; and
- heating the mixture to a temperature above 400C to form AlCl<sub>3</sub> and titanium-aluminium compounds.

In accordance with a fourth aspect, the present

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invention provides an apparatus for the production of a metal compound, comprising:

- a reaction vessel arranged in use for the mixing of aluminium with a precursor material of metal halide;
- the vessel also arranged in use for the resultant mixture to be heated to a temperature sufficient for the precursor material to react with the aluminium to form aluminium halide and a product;
- 5 and
- a first condensation zone arranged in use to operate at a temperature such that the aluminium halide condenses in the first condensation zone.

In one form the apparatus can also comprise a second condensation zone arranged to condense metal halide escaping the reaction mixture. The second condensation zone can be arranged to return condensed metal halide to the reaction zone.

In one embodiment, the reaction zone operates at a temperature T1 and the first condensation zone at a temperature T2 which is lower than the temperature T1. In one form, the second condensation zone operates at a temperature T3 which is between T1 and T2.

In one example, the precursor material may be a material containing titanium as a component, such as titanium trichloride and/or titanium dichloride. Where the precursor material includes titanium trichloride and where the apparatus includes a first condensation zone that operates at temperature T2, T2 is preferably below 200C, wherein gaseous aluminium trichloride emanating from the reaction zone is condensed.

When the apparatus includes a second condensation zone, T3 is below 500C and titanium trichloride which escapes from the reaction zone is recondensed at the second condensation zone. In one embodiment, the second condensation zone is located between the reaction zone and the first condensation zone.

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Typically the apparatus includes a heating arrangement for heating the precursor material. In some embodiments, openings are provided for the introduction of further gases. Openings may also be provided to evacuate the vessel to a low pressure. In still further embodiments, the reaction vessel may be a number of discrete vessels, each vessel providing a different reaction or condensation zone.

Preferably, the apparatus of this aspect of the invention is suitable for implementing the method of any of the foregoing or following aspects of the invention described herein.

In accordance with a fifth aspect, the present invention provides a method for the production of a metal compound, comprising the steps of:

- mixing aluminium with a precursor material of metal halide;
- heating the mixture to a temperature sufficient for the precursor material to react with the aluminium, resulting in the formation of aluminium halide; and
- arranging the conditions to favour a forward reaction to form the aluminium halide and to reduce a reverse reaction to form aluminium and precursor material.

In one form the step of arranging the reaction conditions can include the step of driving aluminium halide away from a reaction zone where the aluminium and precursor material are reacting. Continuous removal of aluminium halide facilitates the forward reaction.

In one embodiment the metal halide can be a titanium subhalide such as titanium trichloride, and a product of the reaction can include titanium compounds. In further embodiments of the method, the precursor material can include a source of one or more elements selected from the group comprising chromium (Cr), niobium (Nb), vanadium (V), zirconium (Zr), silicon (Si), boron (B), molybdenum

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(Mo), tantalum (Ta) and carbon (C).

In accordance with a sixth aspect, the present invention provides a method for the production of vanadium and/or vanadium compounds, comprising the steps of mixing aluminium with a precursor material including vanadium subhalide, and heating the mixture, to form aluminium halides and vanadium and/or vanadium compounds.

The vanadium compounds may include vanadium-aluminium alloys and/or vanadium aluminium intermetallic complexes.

In accordance with a seventh aspect, the present invention provides a method for the production of zirconium and/or zirconium compounds, comprising the steps of mixing aluminium with a precursor material including zirconium subhalide, and heating the mixture, to form aluminium halides and zirconium and/or zirconium compounds.

The zirconium compounds may include zirconium-aluminium alloys and/or zirconium-aluminium intermetallic complexes.

In accordance with an eighth aspect, the present invention provides a method for the production of titanium and/or titanium compounds, comprising the steps of mixing a reducing agent with a precursor material including titanium subhalide, and heating the mixture, to form halides of the reducing agents and titanium and/or titanium compounds. The reducing agent can be selected from the group comprising zinc, magnesium, sodium, aluminium or other like metals.

In accordance with a ninth aspect, the present invention provides a method for production of a powder of titanium-aluminium intermetallic compounds including at least one of  $Ti_3Al$ ,  $TiAl$  and  $TiAl_3$ , and alloys based on titanium-aluminium intermetallics according to the first aspect, wherein starting materials for the method include aluminium powder and at least one of titanium chloride or a titanium subchloride.

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Brief Description of the Drawings

Features and advantages of the present invention will become apparent from the following description of 5 embodiments thereof, by way of example only, with reference to the accompanying drawings, in which:

Figure 1 shows the Gibbs energy of formation of  $\text{AlCl}_{3(g)}$ ,  $\text{TiCl}_3$  and  $\text{TiCl}_3 + \text{Ti-Al}$ ;

10 Figure 2 shows the total Gibbs free energy for reactions leading to formation of Ti-metal based compounds;

Figure 3 illustrates the equilibrium composition of  $\text{TiCl}_4$  - Hydrogen plasma at temperatures of between 300K and 5000K;

15 Figure 4 is a schematic diagram of an apparatus for implementing a process in accordance with an embodiment of the present invention;

20 Figure 5 is a schematic diagram of a further embodiment of an apparatus for implementing a process in accordance with an embodiment of the present invention;

Figure 6 is a schematic diagram illustrating a process for production of titanium based compounds in accordance with an embodiment of the present invention;

25 Figure 7 illustrates the Gibbs free energy for half reactions leading to the formation of titanium-tetrachloride;

Figure 8 is a schematic diagram illustrating a process for production of titanium based compounds in accordance with a further embodiment of the invention;

30 Figure 9 is an XRD spectrum for a Ti-6Al-4V powder produced by an embodiment of the present invention; and

Figure 10 is an XRD spectrum for a Gamma TiAl compound produced by an embodiment of the present invention.

Description of Preferred Embodiments

The following description is of preferred embodiments of processes for producing metal compounds, including fine powder and ingots with specific compositions. The processes are useful for production of forms of metals such as titanium, vanadium and zirconium together with alloys and intermetallic compounds of these metals with a controllable amount of aluminium.

For example, Ti-Al,  $Ti_3Al$ ,  $TiAl_3$ , Ti-Al-Cr and Ti-V-Al can be made with accuracy by varying the aluminium content. The relative amounts of titanium chlorides and aluminium are determined by the required composition of end product. In one embodiment the process comprises the steps of preparing solid metal halides, mixing the halides with aluminium metal and heating the mixture to a temperature  $T_1$  to trigger reactions leading to formation of aluminium chloride at a temperature ( $T_1$ ) above the boiling temperature of aluminium chlorides, and condensing the aluminium chlorides away from the reaction zone at a temperature  $T_2$ , where  $T_2$  is less than  $T_1$ . The driving of the aluminium chloride away from the reaction zone moves the equilibrium of reaction in the forward direction i.e. to formation of aluminium chloride and metal (and other products depending upon reaction conditions and components).

For titanium compounds, titanium subchlorides, (preferably titanium trichloride  $TiCl_3$ ) can be produced from a precursor material of  $TiCl_4$ . The  $TiCl_3$  is mixed with aluminium and then heated to a temperature above 300C so that  $AlCl_3$  is formed in the gas phase and the  $AlCl_3$  is condensed away from the reaction zone at a temperature below 200C, leaving a powder of Ti in the reaction zone containing a percentage of aluminium, as required for the end product.

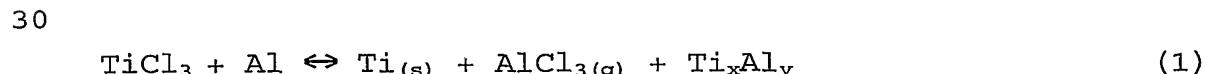
In one embodiment, the process comprises the steps of heating  $TiCl_4$  in a plasma of an argon-hydrogen mixture to

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produce  $TiCl_3$ , and then mixing the resulting  $TiCl_3$  powder with aluminium and then heating the mixture to trigger the reaction. The reaction vessel used is arranged to allow for aluminium chloride to be continuously removed and  
5 condensed in a region away from the reaction zone of the titanium chloride and aluminium mixture. The  $TiCl_3$ , and aluminium in a powder or a lump form (but preferably in a powder form) are mixed together under inert gas or in a vacuum. The mixture is then heated to a temperature of  
10 several hundred degrees to trigger reaction between the two compounds, leading to formation of  $AlCl_{3(g)}$ . The  $AlCl_{3(g)}$  is then condensed elsewhere in the vessel at a temperature below 200C.

In a further embodiment, the process comprises the  
15 steps of heating predetermined amounts of  $TiCl_4$  and aluminium to form  $TiCl_3$  and  $AlCl_3$ , heating the product mixture to a temperature above 300C and providing for  $AlCl_3$  to be evaporated from the reaction zone. The  $AlCl_3$  was driven away from the reaction and condensed away from the  
20 reaction zone at a temperature below 200C. Further aluminium material was then added to the product in an amount depending on the required composition, and then the mixture was heated under the same physical conditions to a temperature above 300C to trigger chemical reactions  
25 leading to formation of  $AlCl_{3(g)}$  whilst providing for the  $AlCl_{3(g)}$  to be condensed elsewhere in the vessel at a temperature below 200C.

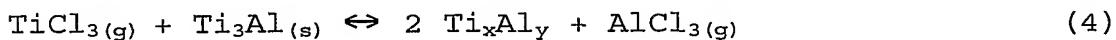
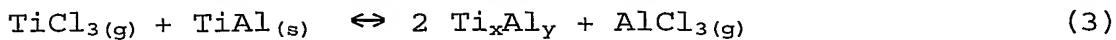
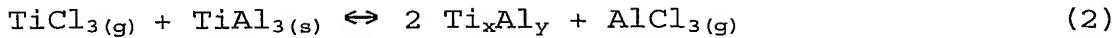
The overall reactions between titanium subchlorides  $TiCl_3$  and Al occur in the following form:



where Al is present in the solid or liquid phase.

The presence of Ti and Al may lead formation of Ti-Al  
35 intermetallic compounds such as  $TiAl_{3(s)}$ ,  $TiAl_{(s)}$  and  $Ti_3Al_{(s)}$ . Then  $TiCl_3$  may react with aluminium according to the following simplified reactions:

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Reactions 1-4 are driven in the forward direction by continuous removal of AlCl<sub>3</sub> from the reaction zone. As a result, equilibrium is moved to the right and the reaction proceeds until completion. The inventor has found that 10 the reaction proceeds slowly at temperatures slightly above 200C under an argon atmosphere at 1 atmosphere. The reaction becomes very rapid at temperatures above 500C as the Gibbs free energy of the total reaction becomes negative as seen in Figures 1 and 2. Figure 1 shows the 15 Gibbs energy for AlCl<sub>3</sub>, TiCl<sub>3</sub>, and TiCl<sub>3</sub>+Ti-Al. Figure 2 shows the total Gibbs energy for Reactions 1-4 leading to formation of solid titanium.

Because of the strong affinity between titanium and aluminium, the presence of Al and Ti may result in 20 formation of titanium-aluminium alloys and/or intermetallic compounds Ti<sub>x</sub>Al<sub>y</sub>. For these compounds, the Gibbs energy of formation ΔG<sub>f</sub> is generally less than 32kJ.mole<sup>-1</sup> for aluminium concentrations up to 80% of the 25 alloys (R.G.Reddy et al. *J. Alloys and Compounds*, vol 321 (2001) 223).

Figure 2 shows the variation with temperature of the total Gibbs energy for reactions leading to formation of AlCl<sub>3(g)</sub> and Ti<sub>(s)</sub>, starting from TiCl<sub>3</sub> and Al. Also shown in Figure 2 is the total Gibbs energy for reaction leading 30 to formation of Ti<sub>(s)</sub> and AlCl<sub>3(g)</sub>, starting from TiCl<sub>3</sub> and Ti-Al compounds. The total Gibbs free energy for Ti-Al is taken to be -32kJ.mole<sup>-1</sup>.

It is usually considered that chemical reactions proceed rapidly for negative values of the total Gibbs 35 energy of the reaction. It is seen in Figure 2 that ΔG is negative at temperatures above 800K (525C) for Reaction 1. This is in excellent agreement with the experimental

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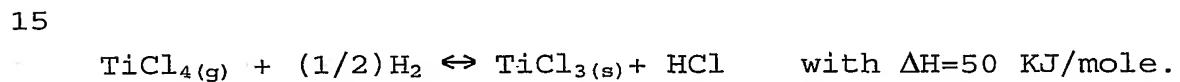
observations which show rapid reaction between  $TiCl_3$  and Al at a temperature of 500C in an argon atmosphere at 1 atmosphere pressure. The inventor found that as the temperature of the mixture  $TiCl_3$ -Al increased above 300C, a 5 cloud of white fume moved from the reaction zone towards the cold region of the vessel where it recondensed to form solid  $AlCl_3$ . At temperatures above 500C, the reaction became almost spontaneous, which is in agreement with the results shown in Figure 2. For reactions involving Ti-Al 10 compounds, the inventor found that in argon at 1 atmosphere pressure, reactions leading to formation of  $Ti_{(s)}$  and  $Ti_xAl_y$  (Reactions 2, 3 and 4) seem to proceed rapidly at temperatures above 850C.

Titanium chlorides may escape from the reaction zone, 15 or disproportionate during heating. Gaseous  $TiCl_3$  that may evaporate during the heating process reacts more readily with Al and further enhances formation of Ti compounds. For a mixture of  $TiCl_3$  and Al powder, with the ratio of  $[Al]/[TiCl_3] > 1$ , the inventor found that only small 20 quantities of less than a few percent of  $TiCl_3$  escape the reaction zone, and are recondensed in a region of the vessel at a temperature around 500K and introduced back into the reaction zone, or alternatively collected for reprocessing. Any  $TiCl_2$  that is produced due to 25 disproportionation, reacts with Al compounds faster than  $TiCl_3$  and enhances reactions, leading to formation of Ti compounds. The inventor found no evidence of major losses due to escape of  $TiCl_4$ . The inventor has made experimental observations which suggest that, for production of 30 titanium with a high aluminium content, disproportionation reactions have little or no significant impact on the efficiency of the process, since most of the Ti in the feedstock materials was able to be accounted for. For production of titanium with a low aluminium content, the 35 initial amount of aluminium used is less than the stoichiometric amount needed to remove all of the chlorine from the  $TiCl_3$  materials. Excess titanium chloride

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remaining after depletion of available aluminium, is evaporated from the product and condensed elsewhere for reprocessing.

Production of  $TiCl_3$  can be carried out from  $TiCl_4$  using a hydrogen plasma route or through reduction with aluminium. Production of  $TiCl_3$  in a hydrogen plasma, known as the Huel process, has been used in industry for several decades. Figure 3 shows the composition of  $TiCl_4$ -Hydrogen plasma at temperature between 300K and 5000K. It is seen that  $TiCl_4$  can be converted into solid  $TiCl_3$  by reacting it with hydrogen in a plasma. It is also seen that the conversion rate is almost 100%. The energy cost for synthesis of solid  $TiCl_3$  is very low as the overall reaction leading to  $TiCl_3$ :



For reduction of  $TiCl_4$  with aluminium, the process is usually carried out in closed vessel containing appropriate amounts of  $TiCl_4$  and Al at a temperature above 200C, leading to formation of a mixture of  $TiCl_3$  and  $AlCl_3$ . Pure  $TiCl_3$  is obtained from the mixture by distillation at temperature above 200C and allowing  $AlCl_3$  to condense elsewhere.

25                   For the process disclosed herein, production of titanium-aluminium compounds is made by mixing titanium subchlorides, preferably  $TiCl_3$ , with aluminium in a powder form, placing the materials in a vessel under vacuum or in an inert atmosphere, and heating the mixture. For 30 processing under flowing inert gas or under vacuum,  $AlCl_3$  formed due to reactions described above is driven into a different part of the vessel at a temperature below 200C. This favours the forward reaction formation of the aluminium halide. The heating continues until the 35 reactions proceed to completion, or until complete depletion of available titanium subchlorides and/or aluminium occurs.

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Figure 4 shows a simple system used to make Ti-Al compounds with different Al contents and compositions. For this configuration, a mixture of  $TiCl_3$  and Al, (1), is placed into a vessel (2) and heated to a temperature 5 higher than 300C (typically up to a temperature of the order of 1000C depending on the composition of the mixture). Reactions between  $TiCl_3$  and Al in the vessel (2) lead to formation of gaseous  $AlCl_3$ . A stream of argon gas 10 (10) that is introduced in the vessel (2) carries the gaseous  $AlCl_3$  together with any titanium chlorides that may escape from the reaction zone and drives them through a second vessel (3) which is held at a temperature between 300C and 500C, so that  $TiCl_3$  is recondensed while  $AlCl_3$  remains in the gas phase. Alternatively,  $TiCl_3$  may be 15 recondensed on the upper walls of the vessel (2) if it is held at an appropriate temperature. The remaining  $AlCl_3$  together with any  $TiCl_4$  that may have formed in the reaction zone due to disproportionation is driven through a vessel (4) at a temperature higher than 136C and lower 20 than 200C so that  $AlCl_3$  is recondensed, and the remaining  $TiCl_4$  is driven into a vessel (5) which is held at room temperature. The remaining argon gas is discharged out of the system or recycled.

Typically the gaseous atmosphere in the vessel is an 25 inert gas, such as argon, helium, neon, xenon. Reactive gases such as methane or oxygen are undesirable as they can chemically react with the mixture resulting in other products. It is noted that the reactions can also be conducted in the absence of a gaseous atmosphere (eg under 30 vacuum).

The  $TiCl_3$  and an aluminium powder, the relative mass of which compared to the mass of  $TiCl_3$  depends on the composition of the required product, are introduced into a vessel as described above and then heated until the 35 reaction is complete.

For these processes described above, the product is typically in the form of a fine powder. The powder may be

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discharged from the vessel, at the completion of chemical reactions in the reaction zone, for further processing.

Alternatively, the powder may be further processed in-situ for production of other materials. Alternatively the

5 powder may be heated in-situ to make coarse grain powder.

In a further embodiment, the powder may be compacted and/or heated in-situ and then melted to produce ingot.

It is highly advantageous to have titanium-aluminium

compounds produced in powder form. As discussed in the

10 preamble this is something that prior art processes cannot do directly. The powder form is much more versatile in manufacture of titanium aluminium alloy products, eg

shaped fan blades that may be used in the aerospace industry.

15 The aluminium to be mixed with the titanium subchloride in these processes is, in one embodiment, in fine powder form, usually having an approximate grain topsize of less than 50 micrometres in diameter. Fine aluminium powder is usually less than 50 micrometres in 20 diameter. A problem with using fine aluminium powder is that it is quite expensive to produce and therefore increases the cost of the process, although the inventor still believes that the cost will still be far less than prior art processes.

25 In an alternative embodiment, coarse aluminium powder is used, the powder having an approximate grain topsize of greater than 50 micrometres in diameter. The coarse aluminium powder is added to the titanium subchloride and the mixture is mechanically milled to reduce the 30 dimensions of the aluminium powder in at least one dimension. This can result in the production of "flakes" of aluminium which have a size in at least one dimension which is less than 50 micrometres and which is sufficient to facilitate a satisfactory reaction between the titanium 35 subchlorides and the aluminium. Flakes provide a higher reaction surface area and the small thickness of the flakes results in a more uniform composition of product.

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In a further alternative embodiment, the aluminium raw material may be obtained in the form of flakes (ie already pre-milled) and mixed with the titanium subchlorides before reaction commences.

5 A further embodiment of an apparatus which can be used to prepare titanium-aluminium compounds in accordance with the present invention is illustrated in Figure 5. The apparatus in this case is a simple vessel (60) having relatively long (tall) side walls (20). An upper portion 10 (40) of the side walls (20) forms a first condensation zone at temperature T2, for condensation of AlCl<sub>3</sub>. A middle portion (50) of the side walls (20) forms a second condensation zone at temperature T3 allowing for condensation of TiCl<sub>3</sub>. Titanium-aluminium compounds (11) 15 are formed at the bottom of the vessel (60).

Parameters influencing reactions in the reaction zone include the pressure in the reaction vessel, the temperature of the reaction zone and the grain size of the Al powder. The inventor has found that, for operation 20 under low pressure, a lower temperature is required in order to drive the reaction, as AlCl<sub>3</sub> is removed faster from the reaction zone and TiCl<sub>3</sub> species become more volatile and more active, thus triggering reactions with aluminium. However, this also results in a lower yield, 25 escape of some volatile titanium chloride, and possibly to the production of a two phase product due to disproportionation.

Also, the inventor found that the reaction between TiCl<sub>3</sub> and Al depends strongly on the size of the Al powder 30 grains. The reaction is much faster for smaller grains and also the yield is higher. Very fine aluminium powder results in the formation of a product of Ti-Al compounds with very fine grains, having irregular shapes. The inventor also found that with cheaper, less fine powders, 35 the production yield of titanium aluminium compounds was still satisfactorily high and the resultant grain size comparable to that achieved with finer aluminium powders.

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As discussed above, relatively coarse titanium powders can also be used, and the mixture can be milled to produce flakes, or the aluminium starting material can be provided in the form of flakes.

5 As discussed above,  $TiCl_4$  can be used to produce the titanium subchlorides to be used as the precursor material for the production of the titanium aluminium compounds. Thus, titanium tetrachloride can be used as a feedstock material. The production of  $TiCl_4$  from titanium ore 10 (titanium oxide) is a well known process, usually as a precursor step for preparation of Ti metal by processes such as the Kroll and Hunter process. Methods in accordance with the present invention can also use  $TiCl_4$  as a feedstock material. Instead of preparing the metal 15 directly from  $TiCl_4$ , however,  $TiCl_4$  is reduced to produce the precursor material  $TiCl_3$ . As briefly described above, this embodiment utilises two methods for the production of  $TiCl_3$ :

20 Reduction of  $TiCl_4$  using Aluminium:

$TiCl_4$  and aluminium metal (coarse or fine powder) in appropriate amounts are introduced into a closed vessel under an inert gas atmosphere (such as argon). The vessel is then heated to a temperature above 200C to form a 25 mixture of  $TiCl_3$  and  $AlCl_3$ . The  $TiCl_3$  powder is then extracted from the mixture by distillation as described before. The  $TiCl_3$  powder is then mixed with more aluminium if required and processed utilising an apparatus such as described above in relation to Figure 4.

30

Reduction of  $TiCl_4$  using hydrogen:

$TiCl_4$  may be fed into a plasma-processing unit operating with argon and hydrogen gas to produce  $TiCl_3$ . Products exiting from the plasma processing system may 35 travel through a filter to separate  $TiCl_3$  from the gas stream and the resulting  $TiCl_3$  powder can then be moved into a processing chamber where it is mixed with an

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appropriate amount of aluminium, depending on the required composition of the end product. The mixture is then processed utilising an apparatus such as that described earlier in relation to Figure 4 or Figure 5. At the 5 completion of the reaction, the materials can be discharged from the reaction vessel for use in manufacturing. Alternatively, the powder can be consolidated in-situ and then melted to produce ingots. Gases from the plasma system may be re-used after 10 separation and cleaning.

In the above-described processes, it is possible to include other precursor materials in addition to the aluminium and titanium subchlorides, to obtain products of desired composition. For example, the precursor materials 15 may include vanadium subchlorides, such as vanadium trichloride and/or vanadium dichloride and the products may include titanium-aluminium-vanadium compounds. The precursor material may include chromium halides and the products may include titanium-aluminium-chromium 20 compounds. Niobium halide may be added as a starter material to produce titanium-aluminium-niobium-chromium compounds. The precursor materials may also include one or more halides of elements such as chromium, niobium, vanadium, zirconium, silicon, boron, molybdenum and 25 carbon.

Figure 6 is a schematic diagram of a process for the production of titanium-aluminium compound powder from titanium tetrachloride starting materials, in accordance with an embodiment of the present invention. The process 30 discloses how aluminium trichloride can be recycled to produce raw materials.

TiCl<sub>4</sub> is reduced using hydrogen, as discussed above, to result in TiCl<sub>3</sub> (Step 1). TiCl<sub>3</sub> is then mixed with aluminium powder and any other precursors that are 35 required are added (Step 2) and then the mixture is processed at temperatures of up to 1000C (temperature will depend upon the precursor mix and products required). Any

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titanium trichloride given off is put back into the reaction mixture (Step 4) and any titanium tetrachloride which is given off is fed back into the process (Step 5) for the production of titanium trichloride (Step 1). From 5 the processing of Step 3, alloy powder products are also obtained (Step 6).

Any aluminium trichloride produced as a by-product (Step 7) can be used for other purposes. For example, such by-products can be electrolysed to produce aluminium 10 and chlorine (the aluminium may be fed back into Step 2). Advantageously, in accordance with an embodiment of the present invention, the aluminium trichloride can be recycled to produce titanium tetrachloride by reacting the  $\text{AlCl}_3$  with the titanium ore (rutile or titanium oxide, Step 15 8; producing titanium tetrachlorides, Step 9; and aluminium oxide, Step 10). The aluminium oxide produced by this process can be sold or electrolysed to produce aluminium raw material, which can be added to the precursor material in this process.

Figure 7 shows the Gibbs free energy for the half reaction leading to aluminium oxide and titanium tetrachloride. The total Gibbs free energy for reaction leading to the formation of titanium tetrachloride is negative at all temperatures higher than 300K, suggesting 25 the reaction is exothermic.

Figure 8 is a schematic diagram of a further production process for the production of titanium aluminium compound powder which involves the step of reduction of titanium tetrachloride with aluminium in 30 order to obtain the required titanium trichloride precursor material. All the other process steps in the Figure 8 production process are the same as the process illustrated in Figure 6 with the exception of Step 1A which is the reduction of titanium tetrachloride by 35 aluminium. Note that Step 1A also may produce some aluminium trichloride by-products which can be recycled via Step 7.

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The following are examples of preparation of titanium aluminium compounds in accordance with an embodiment of the present invention.

5    Example 1: Ti-6Al-4V

TiCl<sub>3</sub> is prepared by reducing TiCl<sub>4</sub> with Al powder. The starting materials were 10g of TiCl<sub>3</sub> + 1g of Al powder (grain size <15 micrometres) +300 mg of VCl<sub>3</sub>. The materials were mixed together very thoroughly and then introduced into a Ta crucible and heated in a quartz tube under flowing argon (100 cc/minute). The temperature is taken to 1000C over 30 minutes and kept there for 1 hour. Materials left in the crucible are 1.65g of metallic powder. The powder is washed in distilled water to remove any residual chlorine (at ppm level) and then dried under argon. XRD analysis of the powder (Figure 9) shows peaks that can be indexed on the Ti-6Al-4V composition. EDX analysis of the powder shows a weight% composition of Ti: 90.1%; Al: 5.8%; V: 4.1%. It was noted that the chlorine and oxygen level were either non existant or below the detection limits of the instrument.

Example 2: Gamma Titanium aluminides

10g of TiCl<sub>3</sub> was mixed with 3.5g of Al powder (grain size <15 micrometres). The mixture was introduced into a Ta crucible and heated in a quartz tube under flowing argon (100 cc/minute). The temperature is taken to 1000C over 30 minutes and kept there for 1 hour. The crucible is then left to cool down and opened. Materials left in the crucible consisted of 4.72 g of grey metallic powder. The powder was washed in distilled water and then dried under argon. XRD analysis (Figure 10) is consistent with the gamma TiAl composition. EDX analysis of the powder suggests the composition of 49.4% (atomic)Ti and 50.6% (atomic)Al.

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Example 3: Ti-48Al-2Cr-2Nb

10g of  $TiCl_3$ , 3.52g of Al powder, 0.34g of  $CrCl_2$  and 0.78g of  $NbCl_5$  were mixed thoroughly and then placed in a Ta crucible in a quartz tube and then heated under flowing 5 argon (100 cc/minute). The temperature was taken to 1000C over a period of 30 minutes and then left at 1000C for 1 hour. 4.4g of metallic powder were left in the crucible. An EDX analysis of the powder suggests a composition of Ti-47Al-2.3Cr-2.3Nb (atomic percent).

10 The methods described herein may also be used for production of metals and metal alloys by mixing metal halide or a mixture of metal halides (chlorides, bromides, iodides and fluorides) and carrying out the process as described above for the  $TiCl_4$ . For example, zirconium and 15 zirconium alloys may be produced using the same procedures described above for Ti and Ti-alloys respectively. For zirconium-based products, the starting material is zirconium chloride. Other examples of metals that can be produced using the present process include vanadium and 20 its alloys and intermetallic compounds. Titanium metal can be produced by the above process following extensive recycling of titanium chlorides. Titanium intermetallic compounds which can be produced include  $Ti_3Al$ ,  $TiAl$  and  $TiAl_3$ . In still further embodiments, reducing agents other 25 than aluminium which may be able to be used with a metal subhalide to produce a metal compound can include zinc, magnesium, sodium, aluminium or other like metals.

The present method may be used for production of 30 powders with a controlled particle size of various compositions including compounds of pure metal, oxides, nitrides of elements such as vanadium and zirconium, as described above for titanium.

Modifications and variations as would be apparent to 35 a skilled addressee are deemed to be within the scope of the present invention.

## CLAIMS

1. A method for the production of titanium compounds, comprising the steps of mixing aluminium with a precursor material including titanium subchloride, and heating the mixture, to form aluminium chlorides and titanium compounds.
2. A method in accordance with claim 1, comprising the step of arranging reaction conditions to favour a forward reaction to form the aluminium chlorides and titanium compounds.
3. A method in accordance with claim 2, wherein the step of arranging the reaction conditions involves driving aluminium chlorides away from a reaction zone where the aluminium and precursor material are reacting.
4. A method in accordance with claim 3, wherein the step of arranging the reaction conditions includes a continuous removal of aluminium chlorides from the reaction zone.
5. A method in accordance with any one of the preceding claims, comprising the step of preparing the titanium subchloride precursor material from titanium chloride ( $TiCl_4$ ).
6. A method in accordance with claim 5, comprising the step of preparing titanium subchloride by reduction of  $TiCl_4$  using aluminium.
7. A method in accordance with claim 5, wherein the step of preparing titanium subchloride is carried out by heating  $TiCl_4$  in a plasma of an argon-hydrogen mixture.
8. A method in accordance with claim 6 or claim 7, comprising the further step of recycling at least some of the aluminium chloride formed, and utilising the aluminium chloride to produce  $TiCl_4$ .
9. A method in accordance with claim 8 wherein the aluminium chloride is used to reduce titanium oxide to produce  $TiCl_4$ .

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10. A method in accordance with claim 9, wherein aluminium oxide is produced by the reduction of titanium oxide and the aluminium oxide is electrolysed to produce aluminium raw material for use in the method of any one of the preceding claims.
- 5 11. A method in accordance with any one of claims 3 to 10, wherein the aluminium chlorides are condensed away from the reaction zone at a temperature lower than that in the reaction zone.
- 10 12. A method in accordance with any one of claims 3 to 11 wherein titanium subchloride which escapes the reaction zone is condensed at a temperature different to that in the reaction zone.
- 15 13. A method in accordance with claim 12, comprising the further step of returning the condensed titanium subchloride to the reaction zone.
14. A method in accordance with any one of the preceding claims, wherein the precursor material includes vanadium subchloride, and a product of said method is an alloy or intermetallic complex including titanium, aluminium and vanadium.
- 20 15. A method in accordance with claim 14, comprising the steps of mixing the precursor material in appropriate proportions and carrying out the method to produce Ti-6Al-4V.
- 25 16. A method in accordance with claim 14, wherein the precursor material includes zirconium subchloride, and a product of the method is an alloy or intermetallic complex including titanium, aluminium, zirconium and vanadium.
- 30 17. A method in accordance with any one of claims 1 to 13, wherein the precursor material includes niobium halide and chromium halide, and a product of said method is an alloy or intermetallic complex including titanium, aluminium, niobium and chromium.
- 35 18. A method in accordance with claim 17, comprising the steps of mixing the precursor materials in

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appropriate proportions and carrying out the method to produce Ti-48Al-2Nb-2Cr.

19. A method in accordance with any one of claims 1 to 13, wherein the precursor material includes a source

5 of one or more elements selected from the group comprising chromium, niobium, vanadium, zirconium, silicon, boron, molybdenum, tantalum and carbon, and products of said method include titanium-aluminium compounds which include one or more of these

10 elements.

20. A method in accordance with claim 19, wherein the source of the element(s) can be a metal halide, a subhalide, a pure element or another compound which includes the element.

15 21. A method in accordance with claim 19 or claim 20, wherein the products also include one or more of an intermetallic compound, a titanium-(selected element)-alloy, and intermediate compounds.

22. A method in accordance with any one of the preceding 20 claims, wherein the aluminium is added in the form of a powder having an approximate upper grain size of less than about 50 micrometres.

23. A method in accordance with any one of claims 1 to 25 21, wherein the aluminium is in the form of a powder of an approximate upper grain size of greater than about 50 micrometres, and the method comprises the step of milling the aluminium powder and titanium subchloride to reduce the grain size of the aluminium powder in at least one dimension.

30 24. A method in accordance with any one of claims 1 to 21, wherein the aluminium is in the form of flakes having a thickness in one dimension of less than about 50 micrometres.

25. A method in accordance with any one of the preceding 35 claims, comprising the further step of adding a reagent to a product of the method to produce a further product.

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26. A method in accordance with any one of the preceding claims, including a pre-processing step of forming the titanium subchloride as precursor material.

27. A method of producing titanium-aluminium compounds, comprising the steps of:

- heating a precursor material of  $TiCl_4$  in a plasma of an argon-hydrogen mixture to produce  $TiCl_3$ ;
- mixing aluminium with said  $TiCl_3$ ; and
- heating the resultant mixture to produce titanium-aluminium compounds and  $AlCl_3$ .

28. A method of producing titanium-aluminium compounds, comprising the steps of:

- heating a mixture of  $TiCl_4$  and aluminium to form  $TiCl_3$  and  $AlCl_3$ ;
- heating the mixture to a reaction zone temperature above 300C and providing for  $AlCl_3$  to be evaporated from the reaction zone;
- adding further aluminium to the mixture; and
- heating the mixture to a temperature above 300C to form  $AlCl_3$  and titanium-aluminium compounds.

29. An apparatus for the production of a metal compound, comprising:

- a reaction vessel arranged in use for the mixing of aluminium with a precursor material of metal halide;
- the vessel also arranged in use for the resultant mixture to be heated to a temperature sufficient for the precursor material to react with the aluminium to form aluminium halide and a product; .  
and
- a first condensation zone arranged in use to operate at a temperature such that the aluminium halide condenses in the first condensation zone.

30. An apparatus in accordance with claim 29, also comprising a second condensation zone arranged to condense metal halide escaping the reaction mixture.

31. An apparatus in accordance with claim 30, wherein the

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second condensation zone is arranged to return condensed metal halide to the reaction zone.

32. A method for the production of a metal compound, comprising the steps of:

5       - mixing aluminium with a precursor material of metal halide;

          - heating the mixture to a temperature sufficient for the precursor material to react with the aluminium, resulting in the formation of aluminium halide; and

10       - arranging the conditions to favour a forward reaction to form the aluminium halide and to reduce a reverse reaction to form aluminium and precursor material.

15 33. A method in accordance with claim 32, wherein the step of arranging the reaction conditions includes the step of driving aluminium halide away from a reaction zone where the aluminium and precursor material are reacting.

20 34. A method in accordance with claim 32 or claim 33, wherein the metal halide is a titanium subhalide, and a product of the reaction includes titanium compounds.

25 35. A method in accordance with any one of claim 32 to claim 34, wherein the precursor material can include a source of one or more elements selected from the group comprising chromium, niobium, vanadium, zirconium, silicon, boron, molybdenum and carbon.

30 36. A method for the production of vanadium and/or vanadium compounds, comprising the steps of mixing aluminium with a precursor material including vanadium subhalide, and heating the mixture, to form aluminium halides and vanadium and/or vanadium compounds.

35 37. A method in accordance with claim 36, wherein the vanadium compounds may include vanadium-aluminium alloys and/or vanadium aluminium intermetallic

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complexes.

38. A method for the production of zirconium and/or zirconium compounds, comprising the steps of mixing aluminium with a precursor material including zirconium subhalide, and heating the mixture, to form aluminium halides and zirconium and/or zirconium compounds.
39. A method in accordance with claim 38, wherein the zirconium compounds may include zirconium-aluminium alloys and/or zirconium-aluminium intermetallic complexes.
40. A method for the production of titanium and/or titanium compounds, comprising the steps of mixing a reducing agent with a precursor material including titanium subhalide, and heating the mixture, to form halides of the reducing agents and titanium and/or titanium compounds.
41. A method in accordance with claim 40, wherein the reducing agent is selected from the group comprising zinc, magnesium, sodium, aluminium or other like metals.
42. A method for production of a powder of titanium-aluminium intermetallic compounds including at least one of  $Ti_3Al$ ,  $TiAl$  and  $TiAl_3$ , and alloys based on titanium-aluminium intermetallics according to any one of claims 1 to 18, wherein starting materials for the method include aluminium powder and at least one of titanium chloride or a titanium subchloride.

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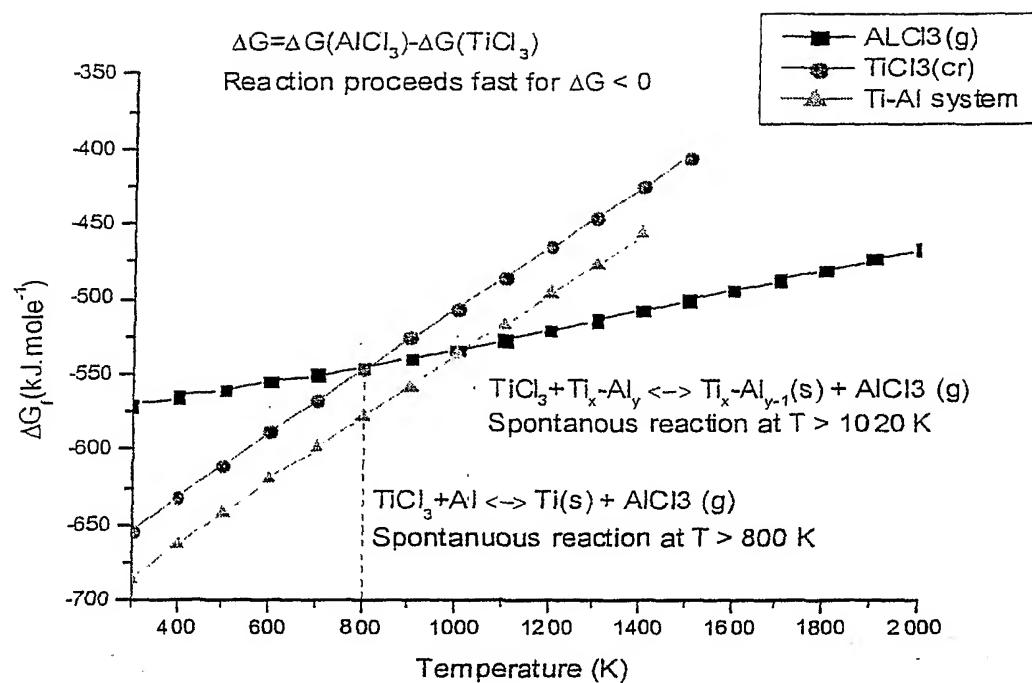


FIG. 1

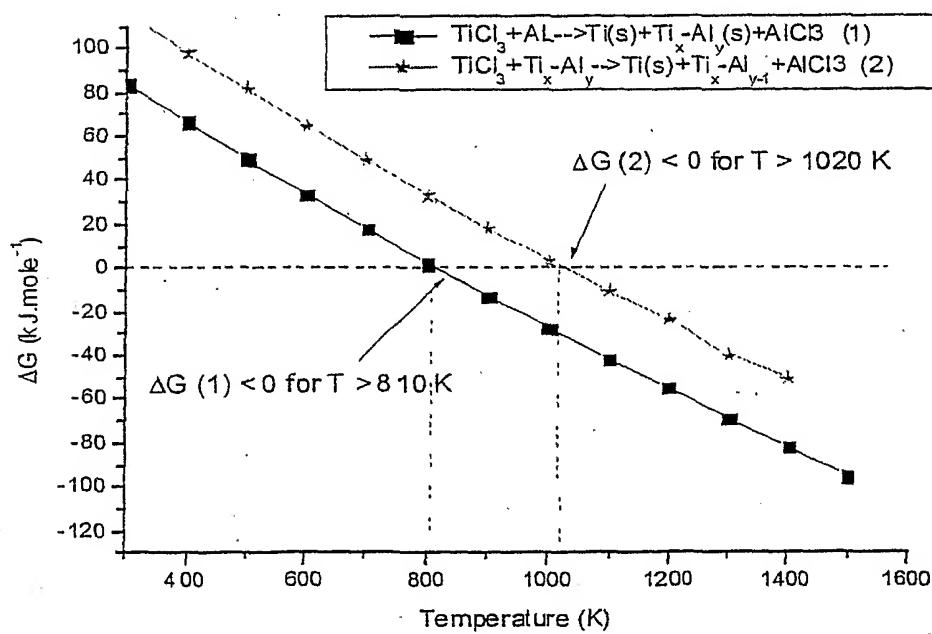


FIG. 2

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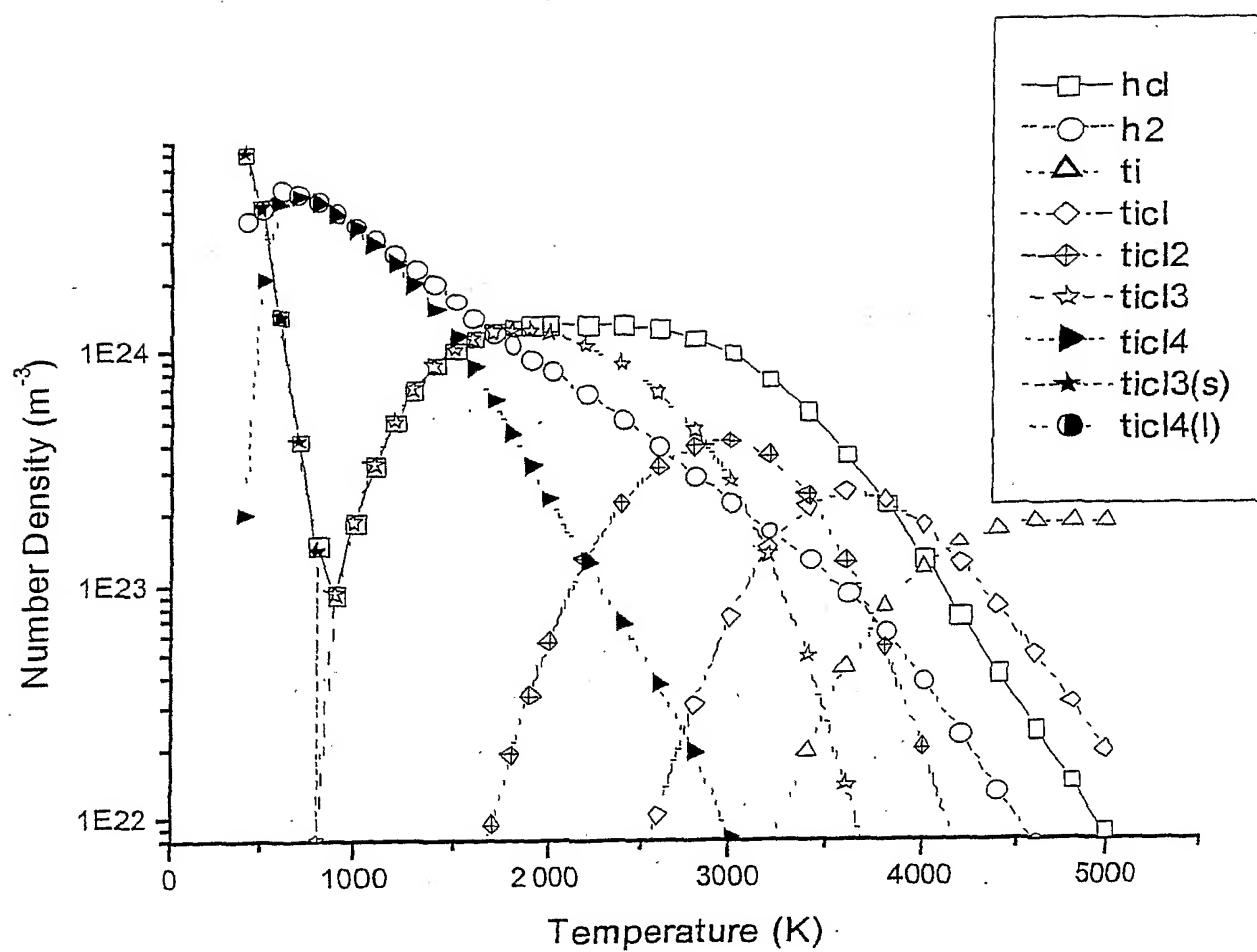


FIG. 3

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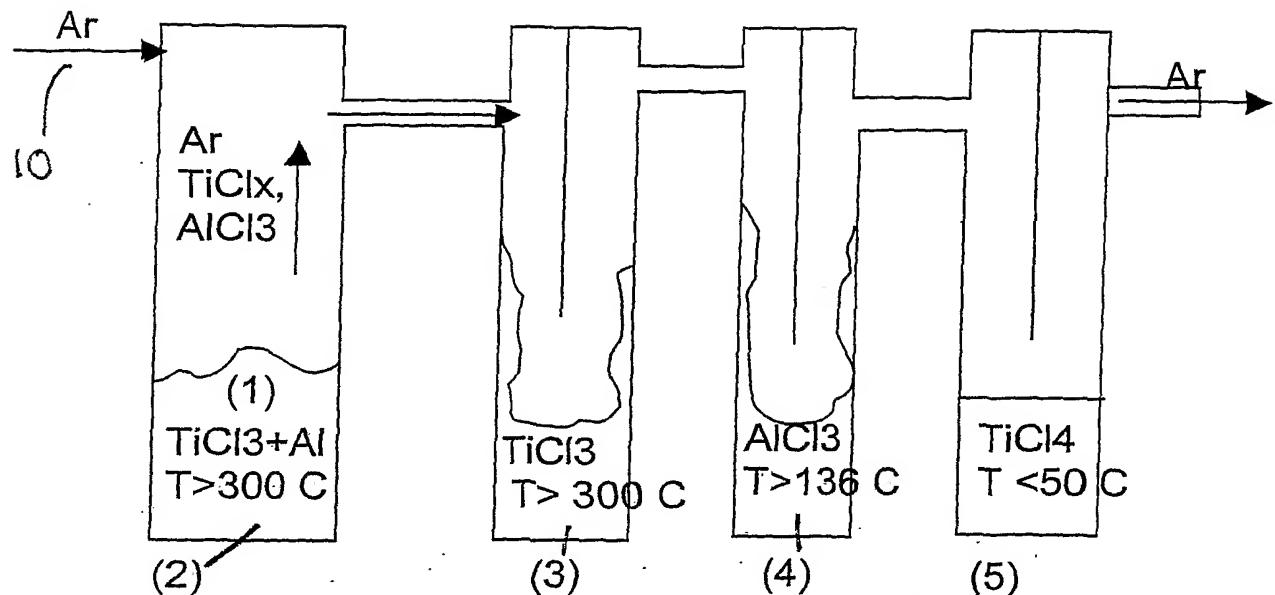


FIG. 4

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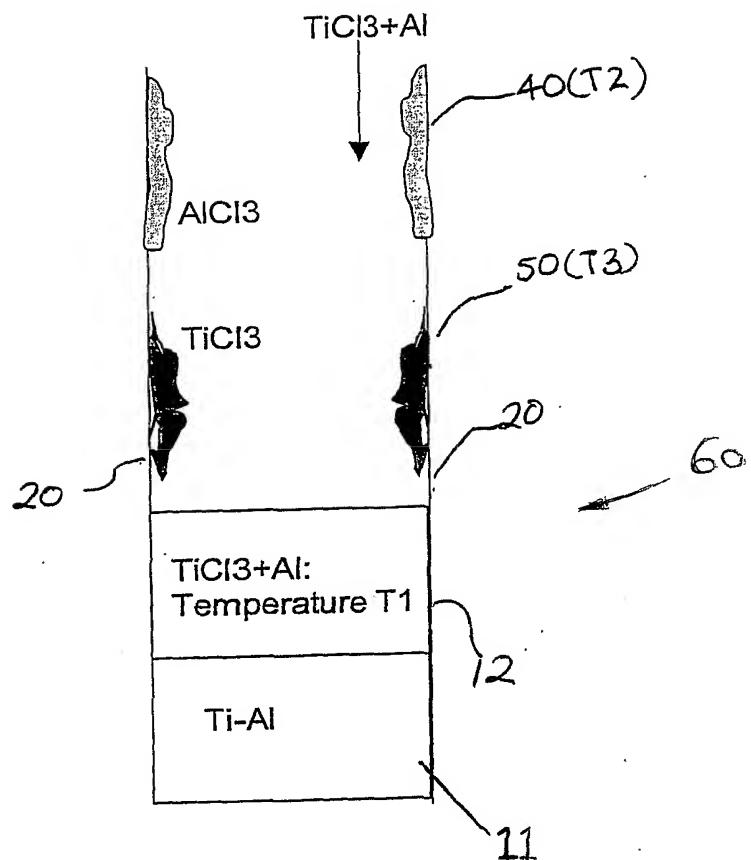


FIG. 5

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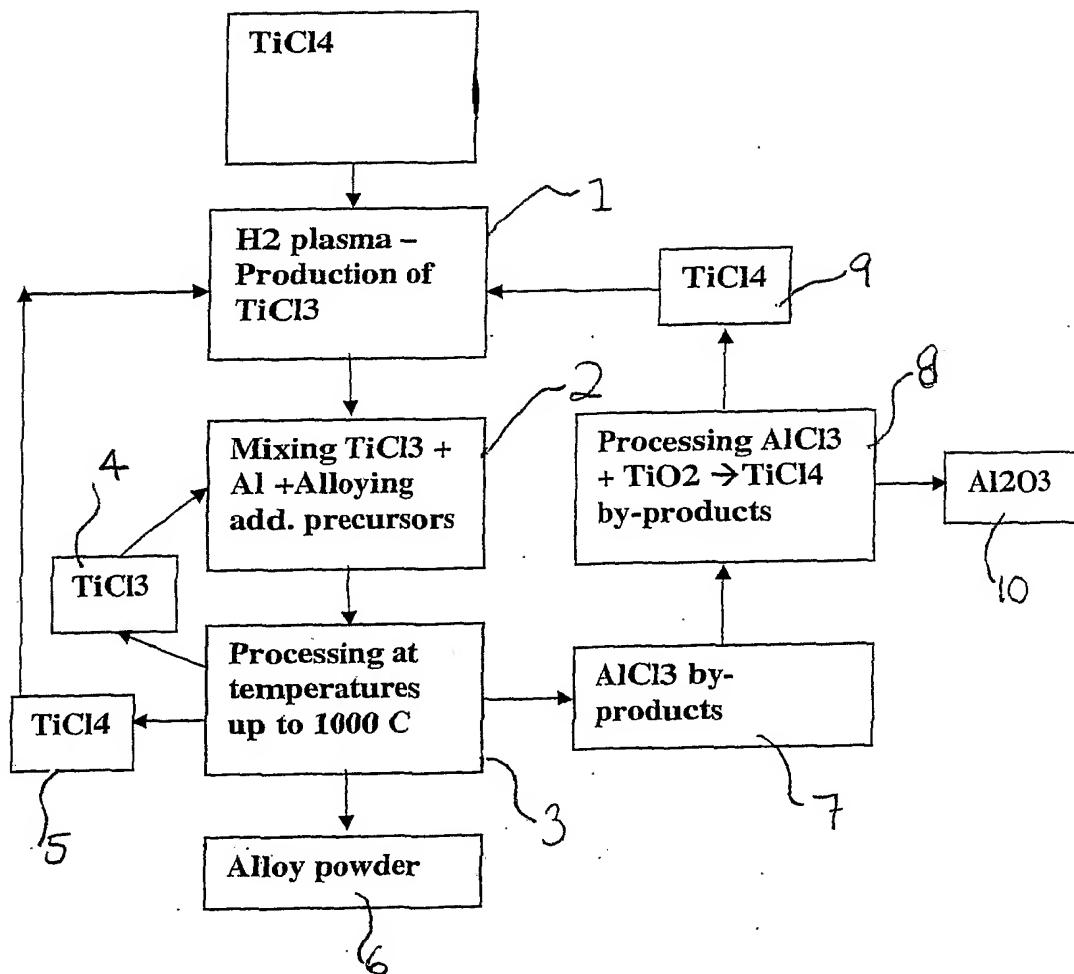


FIG. 6

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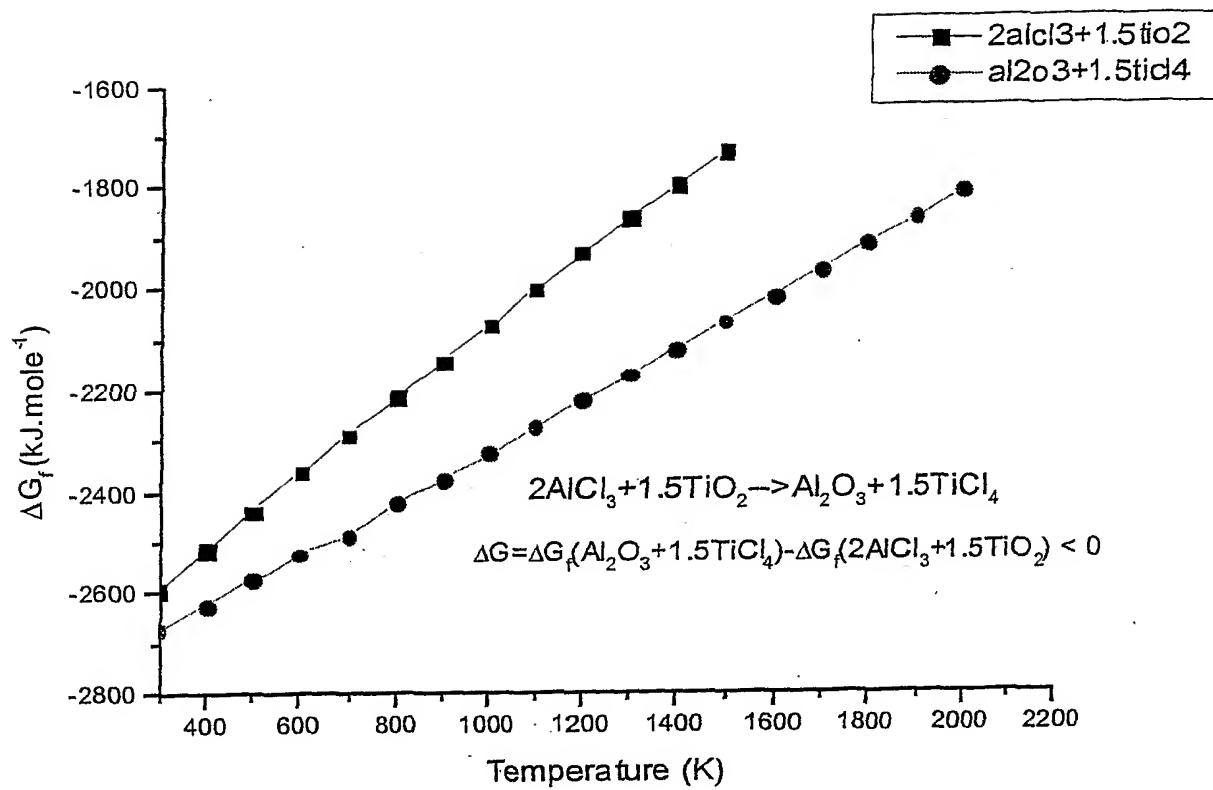


FIG. 7

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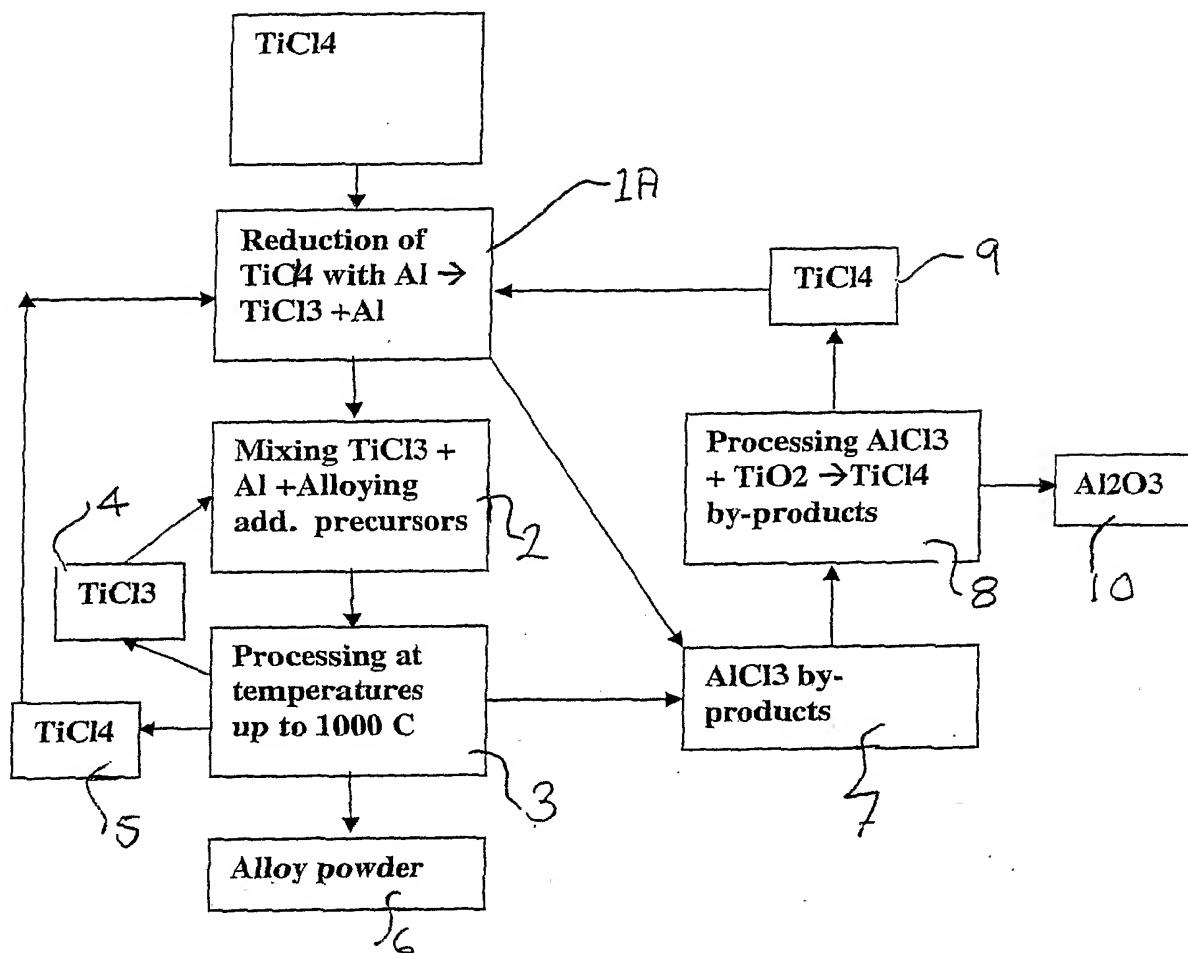


FIG. 8

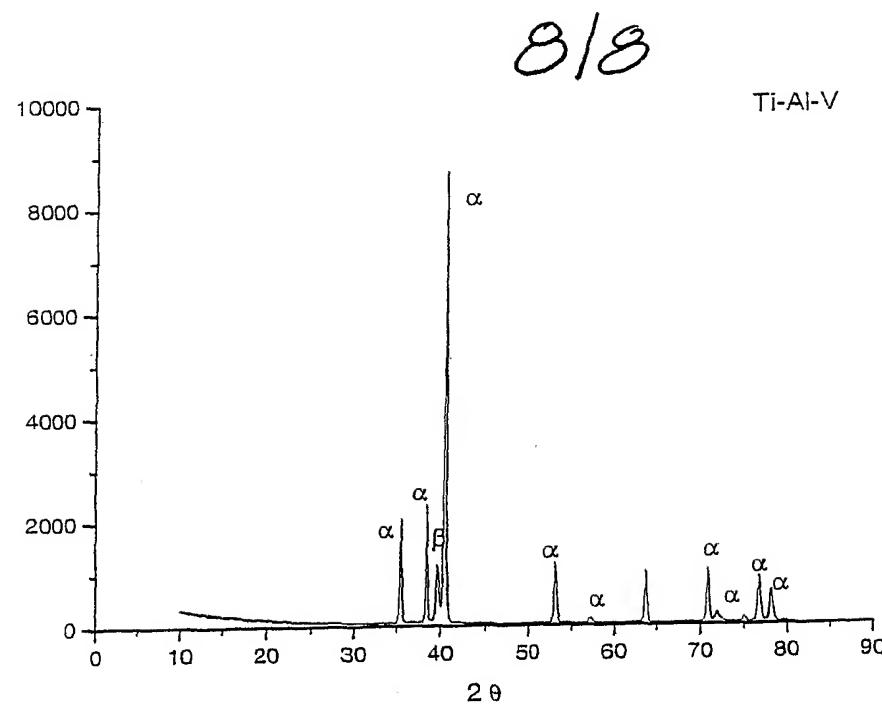


FIG. 9

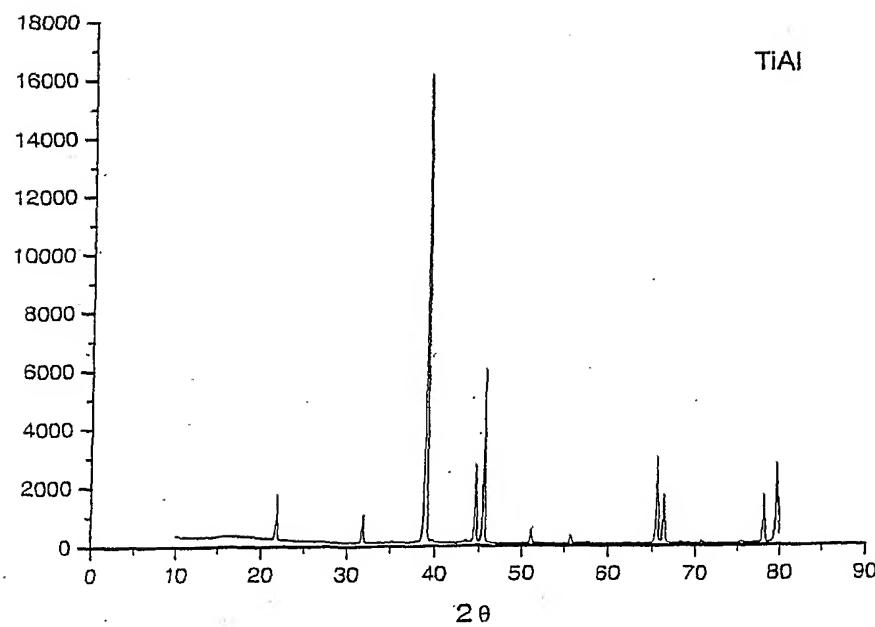


FIG. 10

# INTERNATIONAL SEARCH REPORT

International application No.  
PCT/AU2004/000899

A. CLASSIFICATION OF SUBJECT MATTER																						
Int. Cl. 7: B22F 9/18, 9/20, C22C 14/00, 21/00, 45/08, 45/10, C01F 7/00, C01G 23/00																						
According to International Patent Classification (IPC) or to both national classification and IPC																						
B. FIELDS SEARCHED																						
Minimum documentation searched (classification system followed by classification symbols)																						
Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched																						
Electronic data base consulted during the international search (name of data base and, where practicable, search terms used) DWPI: B22 9/-, C22B 5/04, 34/10, 34/12; C22C 1/00, 1/04, 1/05, 21/-, 14/-; C01F 7/00, 23/00 and keywords for Al/Aluminium, Ti/titanium and +chloride+. Chemical Abstracts Online keywords as above with others for metals/alloys, powders and reduction.																						
C. DOCUMENTS CONSIDERED TO BE RELEVANT																						
Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.																				
X	OKABE, et. al., "Reaction Pathways during Metallothermic Reduction of TiCl <sub>4</sub> by Utilizing Chemical Potential Diagrams," <i>Titanium Extracting and Processing</i> , editors B. Mishra and G. J. Kipouros, The Minerals, Metals and Materials Society, 1997, pp 133-150, See especially pp. 147-150 (note in particular the phase diagram for the Ti subchloride - Al systems on pages 148-149).	1-6, 8-26, 28-42																				
X	Derwent Abstracts Accession No. 58145Y/33, Class E32, JP 52-078608 A (ARITA KENKYUSHO KK) 2 July 1977. See Abstract	1-6, 8-26, 28-42																				
X	Derwent Abstracts Accession No. 98-050427/05, Class M22, RU 2082561 C (TITANIUM MAGNESIUM RESEARCH INST. STOCK. CO.) 27 June 1997. See Abstract	1-6, 8-26, 28-42																				
<input checked="" type="checkbox"/> Further documents are listed in the continuation of Box C		<input checked="" type="checkbox"/> See patent family annex																				
<p>* Special categories of cited documents:</p> <table style="width: 100%; border-collapse: collapse;"> <tr> <td style="width: 20%;">"A"</td> <td>document defining the general state of the art which is not considered to be of particular relevance</td> <td style="width: 20%;">"T"</td> <td>later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention</td> </tr> <tr> <td>"E"</td> <td>earlier application or patent but published on or after the international filing date</td> <td>"X"</td> <td>document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone</td> </tr> <tr> <td>"L"</td> <td>document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)</td> <td>"Y"</td> <td>document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art</td> </tr> <tr> <td>"O"</td> <td>document referring to an oral disclosure, use, exhibition or other means</td> <td>"&amp;"</td> <td>document member of the same patent family</td> </tr> <tr> <td>"P"</td> <td>document published prior to the international filing date but later than the priority date claimed</td> <td></td> <td></td> </tr> </table>			"A"	document defining the general state of the art which is not considered to be of particular relevance	"T"	later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention	"E"	earlier application or patent but published on or after the international filing date	"X"	document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone	"L"	document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)	"Y"	document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art	"O"	document referring to an oral disclosure, use, exhibition or other means	"&"	document member of the same patent family	"P"	document published prior to the international filing date but later than the priority date claimed		
"A"	document defining the general state of the art which is not considered to be of particular relevance	"T"	later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention																			
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"L"	document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)	"Y"	document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art																			
"O"	document referring to an oral disclosure, use, exhibition or other means	"&"	document member of the same patent family																			
"P"	document published prior to the international filing date but later than the priority date claimed																					
Date of the actual completion of the international search 21 July 2004	Date of mailing of the international search report 29 JUL 2004																					
Name and mailing address of the ISA/AU <b>AUSTRALIAN PATENT OFFICE</b> PO BOX 200, WODEN ACT 2606, AUSTRALIA E-mail address: pct@ipaaustralia.gov.au Facsimile No. (02) 6285 3929	Authorized officer <b>ROGER HOWE</b> Telephone No : (02) 6283 2159																					

## INTERNATIONAL SEARCH REPORT

International application No.  
PCT/AU2004/000899

C (Continuation). DOCUMENTS CONSIDERED TO BE RELEVANT		
Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	DeKOCK, et. al., "Attempted Preparation of Ti-6-4 Alloy Powders from $TiCl_4$ , Al, $VCl_4$ and Na," <i>Metallurgical Transactions B</i> , Vol. 18B, September 1987, pp 511-517	1-6, 8-26, 28-42
X	BARTLETT, et. al., "Titanium Powder by Magnesium Reduction of $TiCl_4$ in Liquid Zinc Alloy," <i>Light Metals 1994</i> (edited by U. Mannweiller), The Minerals, Metals and Materials Society, 1994, pp. 1181-1186. See in particular the final paragraphs on page 1186.	1-6, 8-26, 28-42
X	US 5460642 A (LELAND) 24 October 1995. See abstract.	1-6, 8-26, 28-42
X	US 5032176 A (KAMETANI et. al.) 16 July 1991. See column 12 lines 51-68.	1-6, 8-26, 28-42
X	WO 1992/014851 A (THE UNIVERSITY OF MELBOURNE) 3 September 1992.	1-6, 8-26, 28-42
X	EP 134643 A (SOLEX RESEARCH CORPORATION OF JAPAN) 20 March 1985.	1-6, 8-26, 28-42
X	DE 2747016 A (TOYO SODA MANUFACTURING CO. LTD.) 27 April 1978. & Derwent Abstracts Accession No. 32433A/18, Class E32.	1-6, 8-26, 28-42
X	Derwent Abstracts Accession No. 2001-592638/67, Class M22, JP 2001-192711 A (TOHO TITANIUM CO. LTD.) 17 July 2001.	1-6, 8-26, 28-42
X	US 4508322 A (KIMURA et al) 2 April 1985 See Abstract	29-31
X	EP 97135 A (ISHIZUKA) 28 December 1983 See Abstract	29-31
X	GB 1566363 A (VSESOJUZNY NAUCHNO-ISSLEDOVATELSKY I PROEKTNY INSTITUT TITANA et al) 30 April 1980	29-31
X	EP 298698 B1 (TOHO TITANIUM CO. LTD.) 11 January 1989	29-31

**INTERNATIONAL SEARCH REPORT**

International application No.

**PCT/AU2004/000899**

This Annex lists the known "A" publication level patent family members relating to the patent documents cited in the above-mentioned international search report. The Australian Patent Office is in no way liable for these particulars which are merely given for the purpose of information.

Patent Document Cited in Search Report			Patent Family Member							
JP	52078608									
RU	2082561									
US	5460642	AU	21595/95	WO	9525824					
US	5032176	DE	4016502	GB	2231883	JP 2311316				
WO	9214851	AU	12741/92	US	5397375	US 5482691				
		US	5788948							
EP	0134643	AU	30262/84	JP	60017026	JP 60017027				
DE	2747016	JP	53050099							
JP	2001192711									
US	4508322	CA	1218523	FR	2545106	JP 59226127				
		NO	840790							
EP	0097135	AU	14524/83	CA	1202183	JP 58210128				
		JP	59080593	NO	831930	US 4565354				
GB	1566363									
EP	0298698	AU	18824/88	JP	1015334	US 4877445				

Due to data integration issues this family listing may not include 10 digit Australian applications filed since May 2001.

END OF ANNEX